**Enthalpy Definitions**

Match these terms to their definitions.

**The standard enthalpy change of formation (ΔHfӨ)**

**Standard conditions**

**The standard enthalpy change of combustion (ΔHcӨ)**

**The standard enthalpy change of reaction (ΔHrӨ)**

**Standard state**

1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ are a pressure of 100kPa (1 atmosphere), a stated temperature, usually 298 K (25oC) and a concentration of 1.0 mol dm-3 (for reactions with aqueous solutions).
2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is the physical state of a substance under the standard conditions of 100 kPa (1 atmosphere) and 298 K (25 oC)
3. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is the enthalpy change that accompanies a reaction in the molar quantities expressed in a chemical equation under standard conditions, all reactants and products being in their standard states.
4. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ is the enthalpy change that takes place when one mole of a substance reacts completely with oxygen under standard conditions, all reactants and products being in their standard states.
5. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of a compound is the enthalpy change that takes place when one mole of a compound is formed from its constituent elements in their standard states.  
   For an element it is defined as 0 kJ mol-1.